**Atomic characteristics worksheet**

Practice your understanding of the element’s characteristics by filling in the gaps. Tip: use the book, a periodic table or lecture notes.

**Hydrogen (H)**

Atomic number = 1 Atomic mass = 1.008 Valence electrons = 1

Electron configuration: 1s12s2p3s3p3d

**Beryllium (Be)**

Atomic number = 4 Atomic mass = 9.012 Valence electrons = 2

Electron configuration: 1s22s22p3s3p3d

**Carbon (C)**

Atomic number = 6 Atomic mass = 12.01 Valence electrons = 4

Electron configuration: 1s22s22p23s3d3p

**Oxygen (O)**

Atomic number = 8 Atomic mass = 16.00 Valence electrons = 6

Electron configuration: 1s22s22p43s3p3d

**Neon (Ne)**

Atomic number = 10 Atomic mass = 20.18 Valence electrons = 8

Electron configuration:1s22s22p63s3p3d

**Sodium (Na)**

Atomic number = 11 Atomic mass = 22.99 Valence electrons = 1

Electron configuration: 1s22s22p63s13p3d

**Iron (Fe)**

Atomic number = 25 Atomic mass = 55.85 Valence electrons = 2

Electron configuration: 1s22s22p63s23p63d54s2